

Log C- pH Diyagramları

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Log C – pH

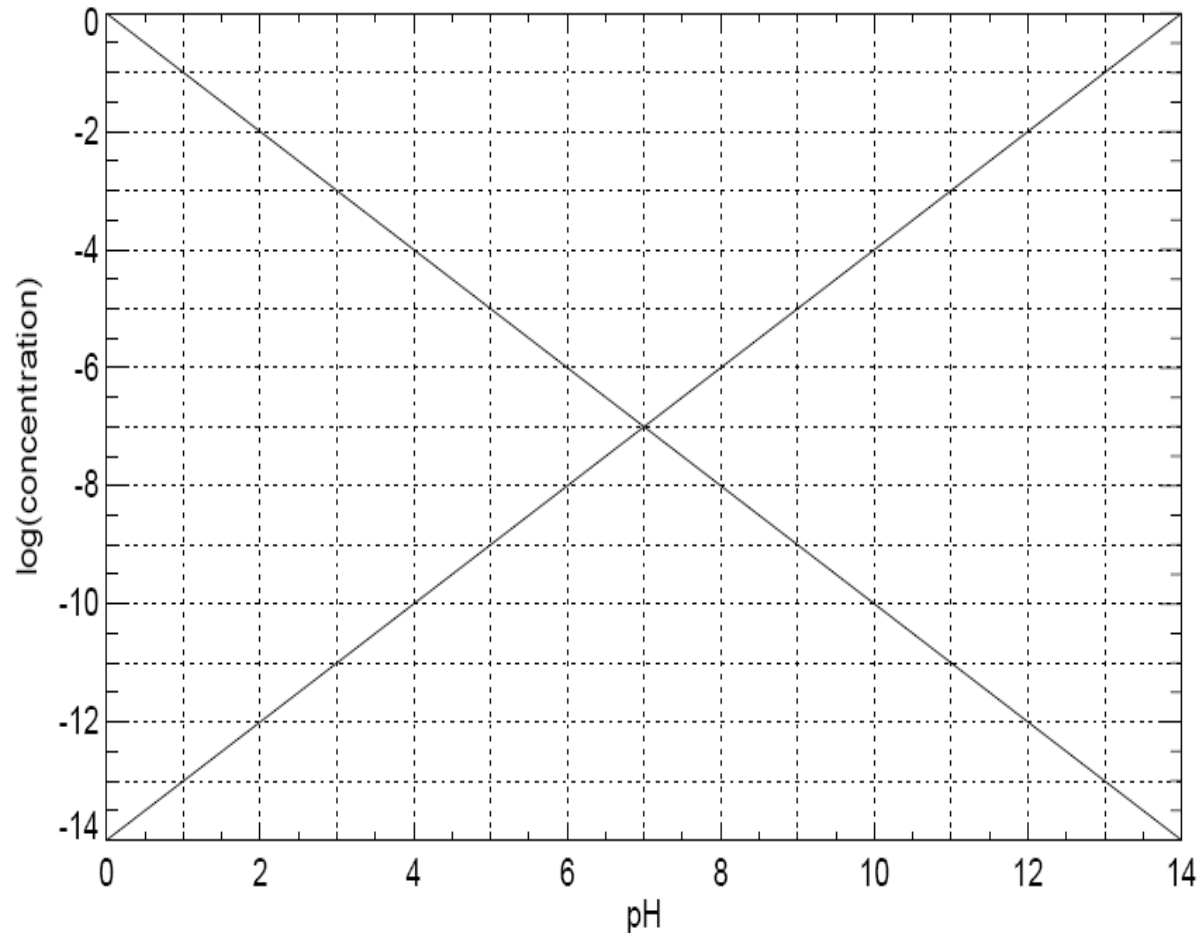
- [] = mol/L
- $[H^+][OH^-]/[H_2O] = K_w = 10^{-14}$
- $[H_2O] = 1$
- $[H^+][OH^-] = 10^{-14}$
- $-\log[H^+] - \log[OH^-] = -\log 10^{-14} = 14$
- $pH + pOH = 14$

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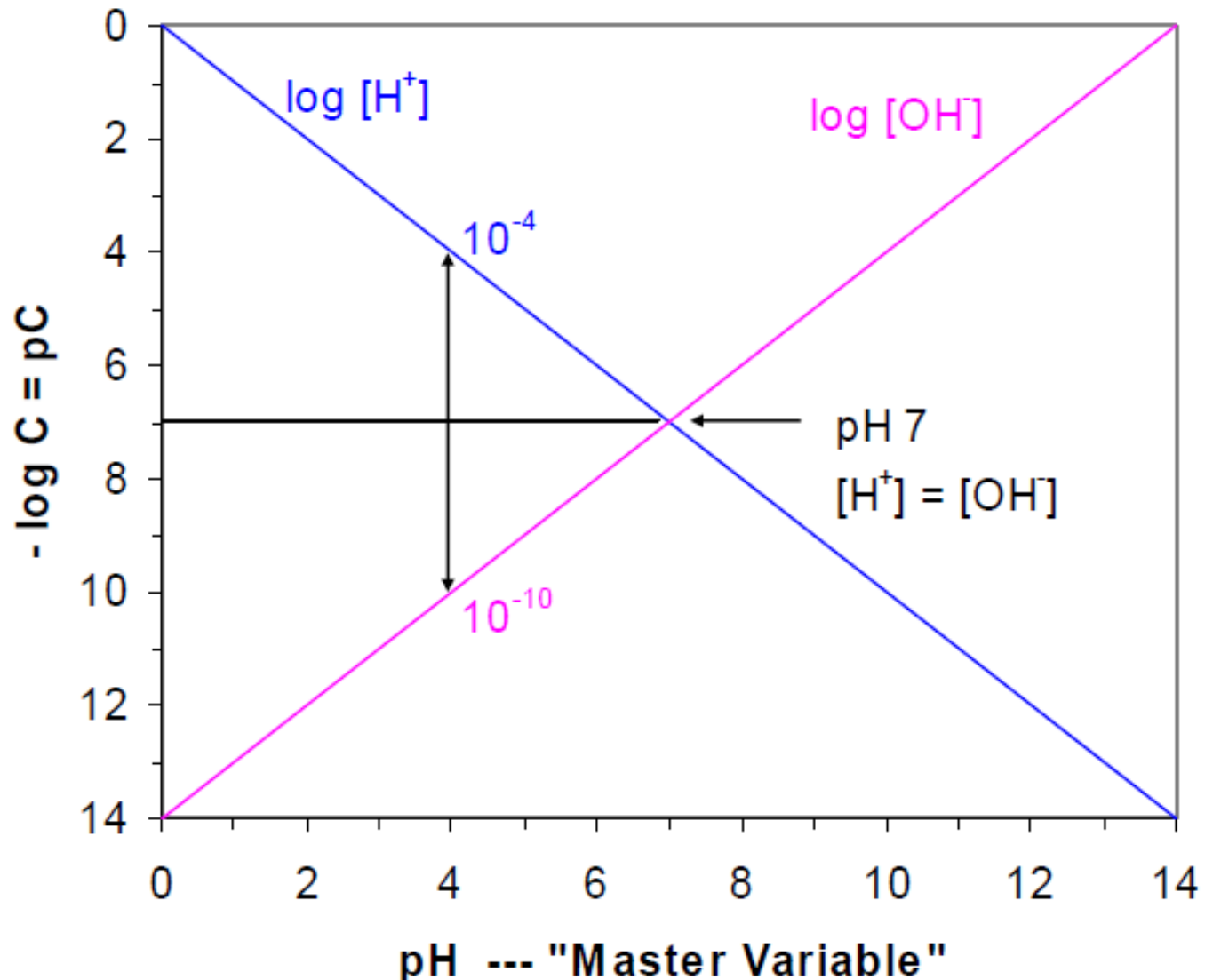
$$[\text{OH}^-] [\text{H}^+] = 10^{-14}$$

$$\underbrace{-\log [\text{OH}^-]}_{-\log C} - \underbrace{\log [\text{H}^+]}_{\text{pH}} = 14$$

Log C - pH

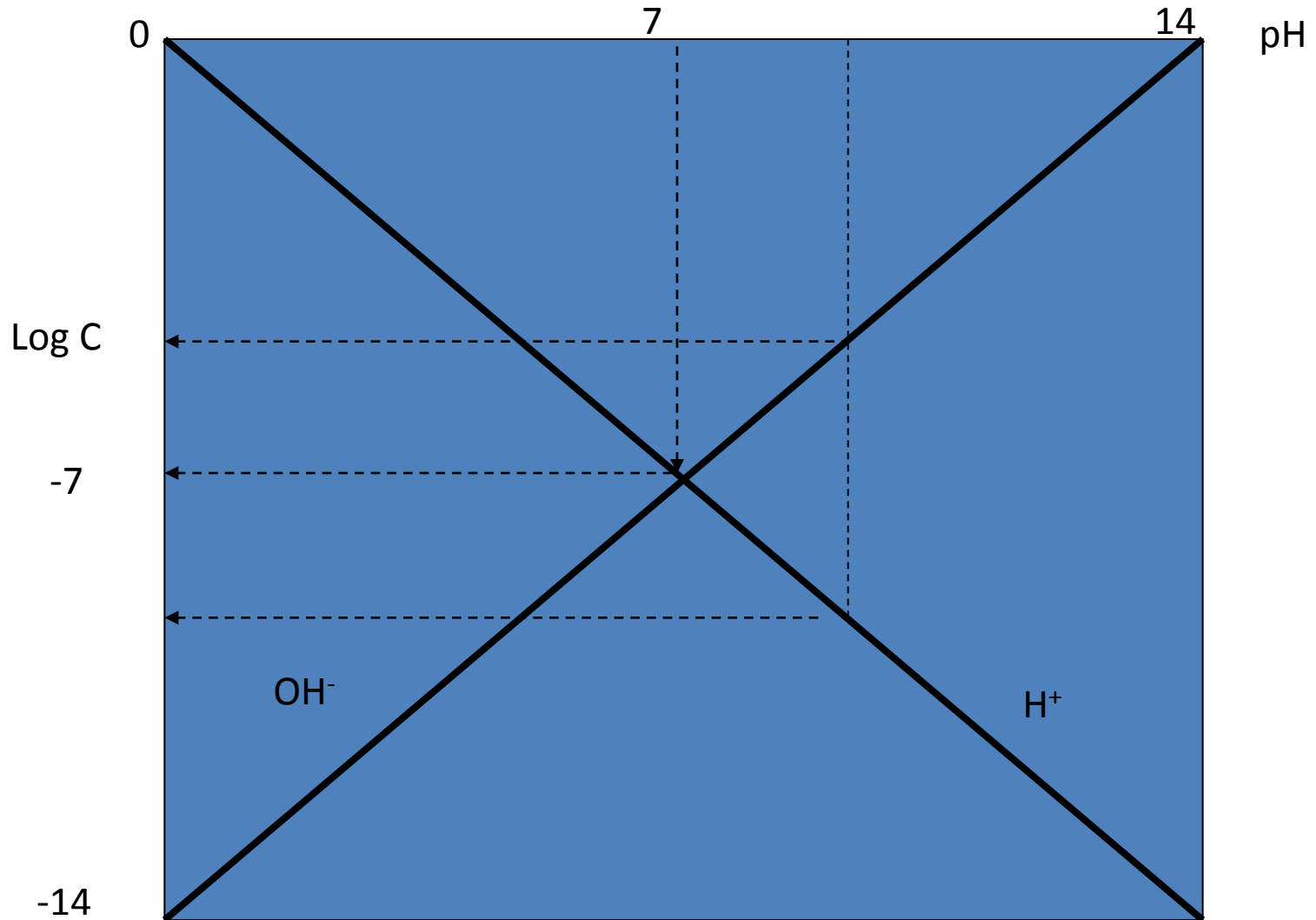


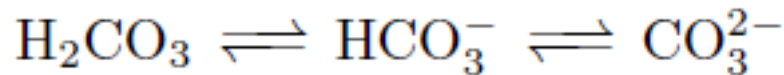
Log C – pH



Log C-pH Diyagramları

$$[\text{OH}^-][\text{H}^+] = 10^{-14}$$





$$\frac{[\text{H}^+][\text{HCO}_3^-]}{[\text{H}_2\text{CO}_3^*]} = K_1 = 10^{-6.3}$$

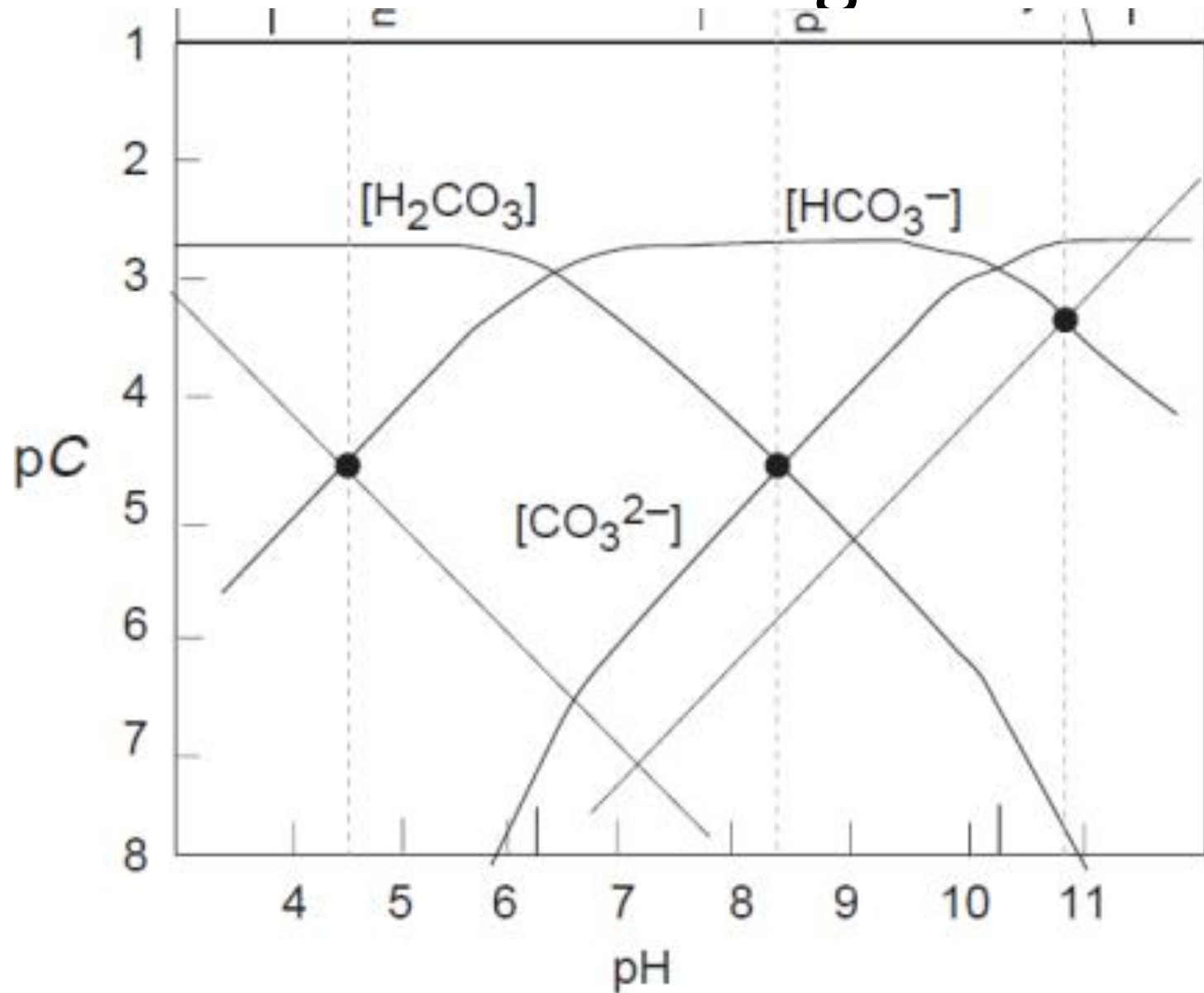
$$\frac{[\text{H}^+][\text{CO}_3^{2-}]}{[\text{HCO}_3^-]} = K_2 = 10^{-10.3}$$

$$[\text{H}^+][\text{OH}^-] = K_w$$

$$C_T = [\text{H}_2\text{CO}_3^*] + [\text{HCO}_3^-] + [\text{CO}_3^{2-}] \quad (\text{mass balance})$$

$$[\text{H}^+] - [\text{HCO}_3^-] - 2[\text{CO}_3^{2-}] - [\text{OH}^-] = 0 \quad (\text{charge balance})$$

Karbonat Dengesi



Problem Example 1

Calculate the pH of a 0.0250 M solution of CO₂ in water.

Solution: Applying the usual approximation $[H^+] = [HCO_3^-]$ (i.e., neglecting the H⁺ produced by the autoprotolysis of water), the equilibrium expression becomes

$$\frac{[H^+]^2}{0.0250 - [H^+]} = 4.47E-7$$

The large initial concentration of H₂CO₃ relative to the value of K_1 justifies the further approximation of dropping the $[H^+]$ term in the denominator.

$$\frac{[H^+]^2}{0.0250} = 4.47E-7$$

$$[H^+] = 1.06E-4; \quad \text{pH} = 3.97$$

temperature	pK_H	pK_1	pK_2	pK_w
<i>fresh water</i> 5 °C	1.19	6.517	10.56	14.73
25	1.47	6.35	10.33	14.00
50	1.72	6.28	10.17	13.26
<i>seawater</i> 25 °C	1.54	5.86	8.95	13.20

Table 3: Some concentration equilibrium constants relating to CO₂ equilibria

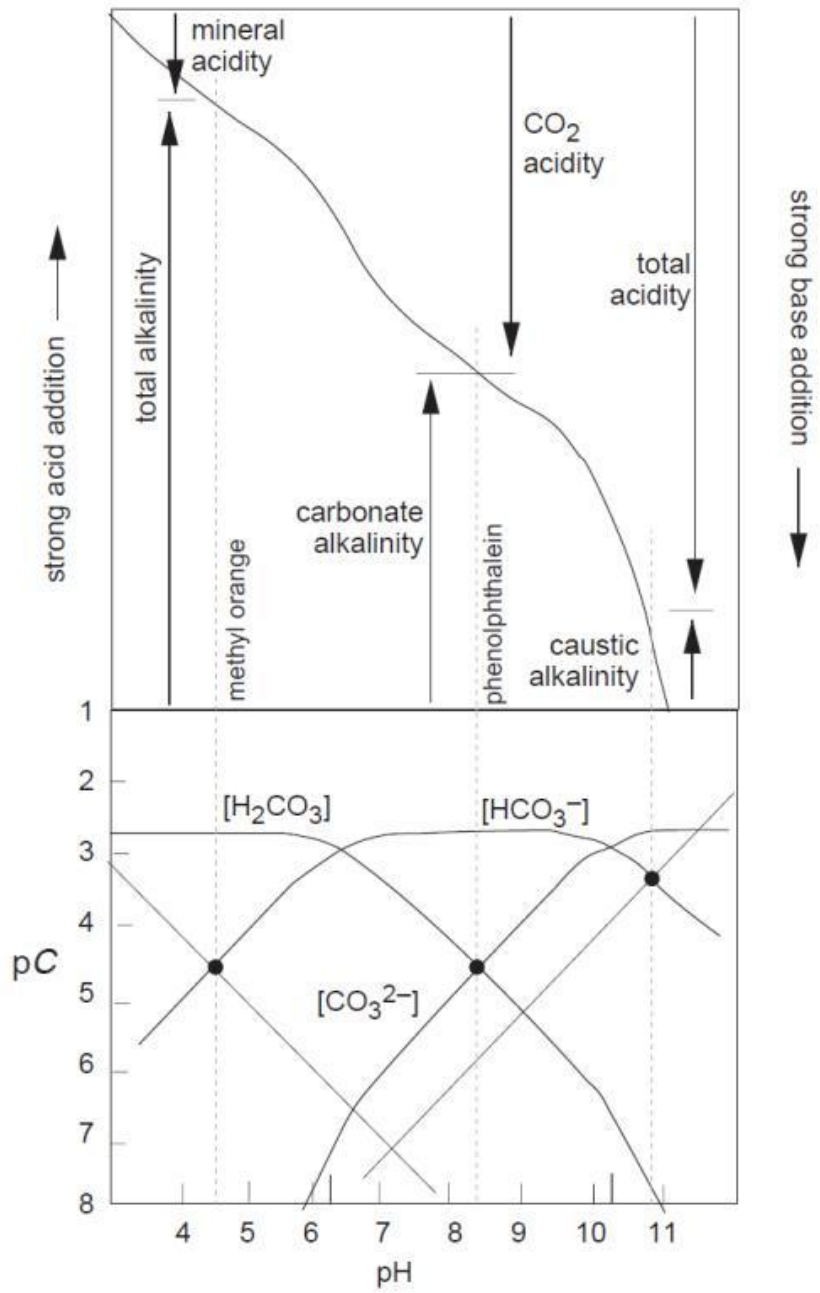


Figure 4: Titration and distribution curves illustrating alkalinity and acidity

Alum Yumaklaşma Diyagramı

Alum Coagulation Diagram

